

Yr 13	Aut 1	Aut 2	Spring 1	Spring 2	Summer 1	Summer 2
	<b>Nomenclature and isomerism &amp; Compound containing the carbonyl group</b>	<b>Aromatic chemistry, amines, polymers</b>	<b>Organic synthesis and analysis &amp; Structure determination</b>	<b>Transition Metals</b>		
	3.3.7 Optical isomerism	3.3.10 Aromatic Chemistry	3.3.14 Organic synthesis	3.2.5.4 Formation of coloured ions		
	3.3.8 Aldehydes and ketones	<b>TEST 3.3.10</b>		3.2.5.5 Variable oxidation states		
	3.3.9 Carboxylic acids and derivatives			3.2.5.6 TM catalysts		
	<b>RP10 Preparation of</b>	<b>Amines</b>	<b>EXAM</b>	<b>RP11 Carry out simple test-tube reactions to identify transition metal ions in aqueous solution</b>		
	<b>1. a pure solid and test of its purity</b>					
	<b>2. a pure organic liquid</b>					
	<b>TEST 3.3.7, 3.3.8 ans 3.3.9</b>	3.3.11 Amines	<b>Organic synthesis and analysis &amp; Structure determination</b>	<b>TEST 3.2.5</b>		
	<b>EXAM</b>	3.3.12 Polymers	3.3.15 Nuclear magnetic resonance spectroscopy			
		<b>Consolidating lesson</b>	<b>RP12 Separating of species by thin-layer chromatography</b>	<b>Reactions of ions in aqueous solution</b>		
			<b>TEST 3.3.14 and 3.3.15</b>	<b>TEST 3.2.6</b>		
		<b>Amino acids, Proteins and DNA</b>	<b>Consolidating lesson</b>			
		3.3.13.2 Proteins		<b>Periodicity</b>		
		3.3.13.3 Enzymes		3.2.4 Properties of Period 3 elements and their oxides		
		3.3.13.4 DNA		<b>TEST 3.2.4</b>		
		3.3.13.5 Action of anti-cancer drugs		<b>Consolidating lesson</b>		
		<b>TEST 3.3.11, 3.3.12 and 3.3.13</b>				
	<b>Kinetics</b>	<b>Acids, bases and buffers</b>	<b>Electrode potentials and electrochemical cells</b>			
	3.1.9.1 Rate equations <b>AND</b> Arrhenius	3.1.12 Acids and bases	3.1.11.1 Electrode potentials and cells			
	3.1.9.2 Determination of rate equation	3.1.12.1 Brønsted–Lowry acid–base equilibria in aqueous	3.1.11.2 Commercial applications of electrochemical cells			
	<b>RP7a and 7b Measuring the rate of reaction by initial rate AND continuous monitoring</b>	3.1.12.3 The ionic product of water, KW	<b>RP8 Measuring the EMF of an electrochemical cell</b>			
	<b>TEST 3.1.9</b>	3.1.12.4 Weak acids and bases Ka for weak acids	<b>TEST 3.1.11</b>			
		3.1.12.5 pH curves, titrations and indicators	<b>Consolidating lesson</b>			
	<b>Equilibrium constant <math>K_p</math></b>	3.1.12.6 Buffer action				
	3.1.10 Equilibrium constant K p for homogeneous systems	3.1.12.3 The ionic product of water, KW				
	<b>TEST 3.1.10</b>	3.1.12.6 Buffer action				
	<b>Consolidating lesson</b>	<b>RP9 Investigate how pH changes when a weak acid reacts with a strong base.</b>				
		<b>TEST 3.1.12</b>				
		<b>Consolidating lesson</b>				

Teacher 1: GVA

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